SC19a Exothermic and endothermic reactions

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| Word | Pronunciation | Meaning |
| **displacement reaction** |  | A reaction where a more reactive element takes the place of a less reactive element in a compound. |
| **endothermic** | end-O-**ther**-mik | A reaction in which energy is transferred to the reactants from the surroundings. The temperature of the surroundings decreases. |
| **exothermic** | ex-O-**ther**-mik | A reaction in which energy is transferred from the reactants to the surroundings. The temperature of the surroundings increases. |
| **neutralisation** | new-tral-i-**zay**-shun | A reaction in which an acid reacts with an alkali or a base to produce a salt and water only. |
| **precipitation** |  | A reaction in which an insoluble product is formed from soluble reactants. |
| **reaction profile** |  | A diagram to show how the energy stored in substances changes during the course of a chemical reaction. |

SC19b Energy changes in reactions

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| Word | Pronunciation | Meaning |
| **activation energy** |  | The minimum amount of energy needed by colliding particles for a reaction to happen. |
| **bond energy** |  | Energy needed to break one mole of a specified covalent bond. |
| **covalent bond** |  | A bond formed when a pair of electrons is shared between two atoms. |
| **mole** |  | A mole of something is 6.02 × 1023 particles of it. The mass of a mole of a substance is the relative formula mass expressed in grams. |
| **reaction profile** |  | A diagram to show how the energy stored in substances changes during the course of a chemical reaction. |